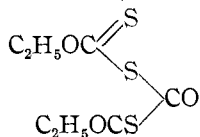


when thiophosgene diluted with acetone is added to an acetone solution of potassium ethylxanthate. The acetone filtrate yields on evaporation a thick, yellowish-brown oil of unknown composition.

*Action of Phosgene, COCl<sub>2</sub>.*—From analogy with the result obtained as described above, it was expected that the diethyl ester of the hypothetical tetrathiotricarbonic acid, C<sub>3</sub>S<sub>3</sub>O(OH)<sub>2</sub>, would be obtained when phosgene in toluene solution was allowed to act on potassium ethylxanthate; however, none of the compound



was obtained. The reaction is apparently complex. Ethylxanthic anhydride in large amount was found in the reaction product.

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*Carborundum and Siloxicon.*—I have read with considerable interest the account of experiments conducted by Messrs. S. A. Tucker and Alexander Lampen on the measurements of temperature in the formation of carborundum, which appeared in this Journal in its July number.

The methods pursued by them were ingenious, but I would say that the temperature they arrived at was somewhat lower than I had anticipated, although I had not made any definite measurements relating to them.

There seems to be, in the text of their article, some confusion regarding siloxicon. This word I coined as a name of a silicon-oxygen-carbon compound and not for amorphous silicon carbide, as stated by them. The zone of material immediately surrounding carborundum (crystalline carbide of silicon) is mainly or wholly composed of an oxygen compound with carbon as an impurity, this carbon being eventually utilized for the extraction of the residual oxygen and formation of a definite silicon carbide. While this apparent inaccuracy does not in any way detract from the interest and value of their determinations, I can scarcely let such a positive statement of the chemical composition of siloxicon pass unnoticed.

EDWARD GOODRICH ACHESON.

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